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## **Chapter 21 Nuclear Chemistry Section**

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Chapter 21. Nuclear Chemistry. 21.2 Nuclear Equations. Learning Objectives. By the end of this section, you will be able to: Identify common particles and energies involved in nuclear reactions; Write and balance nuclear equations; Changes of nuclei that result in changes in their atomic numbers, ...

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## **21.2 Nuclear Equations - Chemistry**

Nuclear Chemistry Nuclear Transformations • Rutherford in 1919 performed the first nuclear transformation. • The transmutations are sometimes represented by listing in order, the target nucleus, the bombarding particle, the ejecting particle and the product nucleus. • The

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above equation becomes:  ${}_{14}^{27}\text{N} + {}_2^4\text{He} \rightarrow {}_{17}^{31}\text{O} + {}_1^4\text{H}$

### **Chapter 21 Nuclear Chemistry**

Nuclear chemistry is the study of reactions that involve changes in nuclear structure. The chapter on atoms, molecules, and ions introduced the basic idea of nuclear structure, that the



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nucleus of an atom is composed of protons and, with the exception of  ${}^1_1\text{H}$ , neutrons.

## **21.1 Nuclear Structure and Stability - Chemistry**

Nuclear Chemistry Section 1 The

Nucleus Section 2 Radioactive Decay

Section 3 Nuclear Radiation Section 4

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Nuclear Fission and Nuclear Fusion

CHAPTER 21. ... nuclear binding and decay. 642 Chapter 21. C A B Number of protons Number of neutrons 0 0 10 10 20 30 40  $n/p = 1.5:1$   $n/p = 1:1$  50 60 70 80 90 100 20 30 40 50 60 70 80 90 100 110 120 130

## **CHAPTER 21 Nuclear hestr - Regents**

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## **Chemistry**

21.1 Nuclear Structure and Stability - Chemistry. Nuclear chemistry is the study of reactions that involve changes in nuclear structure. The chapter on atoms, molecules, and ions introduced the basic idea of nuclear structure, that the nucleus of an atom is composed of protons and, with the exception of ,

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neutrons. Recall that the number of protons in the nucleus is called the atomic number ...

## **Chapter 21 Nuclear Chemistry Answer Key**

np+0e-1 1 1 1 0! pn+0e 1 1 0 1 1! pe1n  
0 0 1 + "1! 4! 2 AP Chemistry CHAPTER  
21- Nuclear Chemistry 21.1 Radioactivity

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- When nuclei change spontaneously, emitting energy, they are said to be radioactive.
- Nuclear chemistry is the study of nuclear reactions and their uses.
- Nucleons are particles in the nucleus: •p<sup>+</sup>: proton •n<sup>0</sup>: neutron
- Atomic number is the number of p<sup>+</sup>.

### **AP Chemistry CHAPTER 21- Nuclear**

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## **Chemistry**

Modern Chemistry 175 Nuclearchemistry  
CHAPTER 21 REVIEW Nuclear Chemistry  
SECTION 4 SHORT ANSWER Answer the  
following questions in the space  
provided. 1. Match each of the following  
statements with the process(es) to  
which they apply, using one of the  
choices below: (1) fission only (3) both

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fission and fusion

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Section Functional Groups Section The  
Nucleus 1 3 5 7 9 b d c a d 2 4 6 8 10 a  
b. Answer Key Nuclear A and B ...

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Nuclear fission and Fusion, Demonstrate  
Equations that Represent the Changes  
that Occur During Radioactive Decay,

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A nuclear fuel. A fissionable isotope must be present in large enough

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quantities to sustain a controlled chain reaction. The radioactive isotope is contained in tubes called fuel rods. A moderator. A moderator slows neutrons produced by nuclear reactions so that they can be absorbed by the fuel and cause additional nuclear reactions. A coolant.



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CHAPTER 21 REVIEW Nuclear Chemistry  
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following questions in the space  
provided. 1. \_\_\_\_\_ The nuclear equation  
is an example of an equation that  
represents (a) alpha emission. (b) beta

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emission. (c) positron emission. (d) electron capture. 2.

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